

Chemistry

Standard level

Paper 1

2 November 2023

Zone A morning | Zone B morning | Zone C morning

45 minutes

Instructions to candidates

- Do not open this examination paper until instructed to do so.
- Answer all the questions.
- For each question, choose the answer you consider to be the best and indicate your choice on the answer sheet provided.
- The periodic table is provided for reference on page 2 of this examination paper.
- The maximum mark for this examination paper is **[30 marks]**.

The Periodic Table

	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1	1 H 1.01																	2 He 4.00
2	3 Li 6.94	4 Be 9.01															9 F 19.00	10 Ne 20.18
3	11 Na 22.99	12 Mg 24.31															17 Cl 35.45	18 Ar 39.95
4	19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.87	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.69	29 Cu 63.55	30 Zn 65.38	31 Ga 69.72	32 Ge 72.63	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.90
5	37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.96	43 Tc (98)	44 Ru 101.07	45 Rh 102.91	46 Pd 106.42	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.60	53 I 126.90	54 Xe 131.29
6	55 Cs 132.91	56 Ba 137.33	57 † La 138.91	72 Hf 178.49	73 Ta 180.95	74 W 183.84	75 Re 186.21	76 Os 190.23	77 Ir 192.22	78 Pt 195.08	79 Au 196.97	80 Hg 200.59	81 Tl 204.38	82 Pb 207.2	83 Bi 208.98	84 Po (209)	85 At (210)	86 Rn (222)
7	87 Fr (223)	88 Ra (226)	89 ‡ Ac (227)	104 Rf (267)	105 Db (268)	106 Sg (269)	107 Bh (270)	108 Hs (269)	109 Mt (278)	110 Ds (281)	111 Rg (281)	112 Cn (285)	113 Unt (286)	114 Uug (289)	115 Uup (288)	116 Uuh (293)	117 Uus (294)	118 Uuo (294)

Atomic number
Element
Relative atomic mass

†

58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	61 Pm (145)	62 Sm 150.36	63 Eu 151.96	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	69 Tm 168.93	70 Yb 173.05	71 Lu 174.97
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90 Th 232.04	91 Pa 231.04	92 U 238.03	93 Np (237)	94 Pu (244)	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf (251)	99 Es (252)	100 Fm (257)	101 Md (258)	102 No (259)	103 Lr (262)
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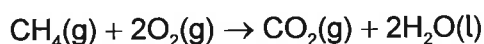
1. Which hydrocarbon would produce equal numbers of moles of CO_2 and H_2O upon complete combustion?

A. CH_4
 B. C_2H_2
 C. C_3H_8
 D. C_4H_8

2. Metal M reacts with 16.0 g of sulfur to produce 26.0 g of the compound MS_2 . What is the relative atomic mass of M?

A. 5
 B. 10
 C. 20
 D. 40

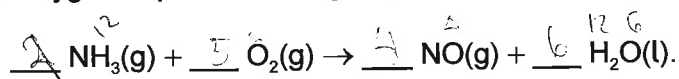
3. 64 g of methane and 96 g of oxygen are reacted according to the equation.



What would be found in the reaction vessel at completion of the reaction?

A. $\text{CO}_2(\text{g})$ and $\text{H}_2\text{O}(\text{l})$ only
 B. $\text{O}_2(\text{g})$, $\text{CO}_2(\text{g})$ and $\text{H}_2\text{O}(\text{l})$ only
 C. $\text{CH}_4(\text{g})$, $\text{CO}_2(\text{g})$ and $\text{H}_2\text{O}(\text{l})$ only
 D. $\text{CH}_4(\text{g})$, $\text{O}_2(\text{g})$, $\text{CO}_2(\text{g})$ and $\text{H}_2\text{O}(\text{l})$

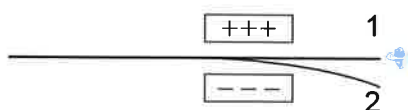
4. Ammonia reacts with oxygen to produce nitrogen (II) oxide and water.



What is the $\text{NH}_3:\text{O}_2$ ratio in the balanced equation?

A. 2:5
 B. 4:5
 C. 1:1
 D. 2:1

5. Gallium ($A_r = 69.72$) consists of two stable isotopes, Ga-69 and Ga-71. What is the relative abundance of Ga-71?
- A. 36 %
B. 40 %
C. 60 %
D. 64 %
6. Which electron configuration represents a d-block element in the ground state?
- A. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^1$
B. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^{10}$
C. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^9$
D. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^{10} 4p^1$
7. A beam containing two different kinds of particles is passed through oppositely charged plates with the results shown in the diagram.



What conclusion can be drawn from this observation?

- A. Particle 1 has a larger mass than particle 2.
B. Particle 2 has a larger mass than particle 1.
C. Particle 1 is positively charged.
D. Particle 2 is positively charged.
8. Which factor generally increases when first ionization energy increases?
- A. Atomic radius
B. Electronegativity
C. Metallic character
D. Nuclear charge

9. The periodic table provided shows 118 elements. Which group of elements would a new element with atomic number 119 be most similar to?
- A. Alkali metals
 - B. Halogens
 - C. Lanthanoids and actinoids
 - D. Noble gases
10. The formula for praseodymium phosphate is PrPO_4 . What is the formula for praseodymium oxide?
- A. Pr_2O
 - B. PrO
 - C. Pr_2O_3
 - D. PrO_2
11. For which molecule can resonance structures be used to describe the bonding?
- A. HCN
 - B. H_2CO_3
 - C. PCl_3
 - D. SO_2
12. What are the electron domain and molecular geometries of SO_3 ?

	Electron domain geometry	Molecular geometry
A.	Trigonal planar	Trigonal planar
B.	Trigonal planar	Trigonal pyramidal
C.	Trigonal pyramidal	Trigonal planar
D.	Trigonal pyramidal	Trigonal pyramidal

13. Which substance has high volatility in its pure state **and** high electrical conductivity in aqueous solutions?

- A. $\text{C}_6\text{H}_5\text{Cl}$
- B. HCl
- C. NaCl
- D. HCN

14. Which reactions release heat?

- I. $\text{C(s)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)}$
- II. $\text{Na}^+\text{(g)} + \text{e}^- \rightarrow \text{Na(g)}$
- III. $\text{NH}_3\text{(g)} \rightarrow \text{NH}_3\text{(l)}$

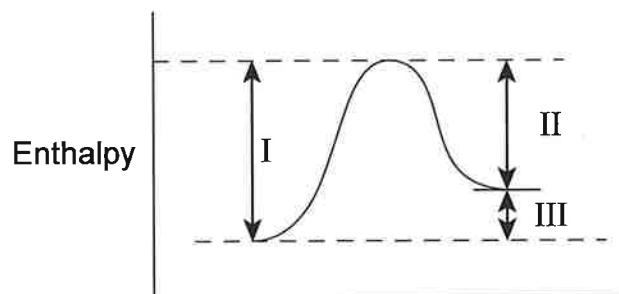
- A. I and II only
- B. I and III only
- C. II and III only
- D. I, II and III

15. Which expression represents the calculation used to obtain the ΔH^\ominus value for the conversion of oxygen to one mole of ozone (O_3)?

		$\Delta H^\ominus, \text{kJ}$
Eqn (i)	$2\text{CO}_2 \rightarrow 2\text{CO} + \text{O}_2$	+566
Eqn (ii)	$3\text{CO} + \text{O}_3 \rightarrow 3\text{CO}_2$	–992

- A. $-566 - 992$
- B. $-566 + 992$
- C. $1.5 \times (-566) + 992$
- D. $1.5 \times (-566) - 992$

16. Which expression represents the calculation of ΔH ?



- A. I-II
 B. II-I
 C. I-III
 D. II-III
17. Which statement describes a role that a catalyst might have in increasing the rate of reaction by providing an alternative mechanism?
- A. It increases frequency of collisions between molecules
 B. It increases energy of collisions between molecules
 C. It increases proportion of molecules colliding in correct orientation
 D. It increases proportion of molecules with a given energy
18. Which of the following equilibria would shift left with an increase in pressure?
- A. $\text{H}_2(\text{g}) + \text{Br}_2(\text{g}) \rightleftharpoons 2\text{HBr}(\text{g})$
 B. $\text{C}(\text{s}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + \text{H}_2(\text{g})$
 C. $\text{NO}(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{NO}_2(\text{g})$
 D. $4\text{NH}_3(\text{g}) + 3\text{O}_2(\text{g}) \rightleftharpoons 2\text{N}_2(\text{g}) + 6\text{H}_2\text{O}(\text{l})$
19. What is the correct way to express the formula for the conjugate base of ethanoic acid?
- A. $^-\text{CH}_3\text{COO}$
 B. CH_3COO^-
 C. $\text{C}_2\text{H}_3\text{O}_2^-$
 D. $^-\text{C}_2\text{H}_3\text{O}_2$

20. ^{SO₂} Sulfur dioxide emissions from coal-fired power plants is a source of acid deposition. Which are pre-combustion methods of reducing sulfur dioxide emissions?
- I. Wash flue gases with crushed limestone and water.
 - II. Crush and wash the coal.
 - III. Crush and mix coal with a sulfur solvent, then wash.
- A. I and II only
 - B. I and III only
 - C. II and III only
 - D. I, II and III
21. In which compound does vanadium have an oxidation state of +4?
- A. $V(NO_3)_2$
 - B. $V(SO_4)_2$
 - C. $V_3(PO_4)_5$
 - D. $V_3(PO_4)_2$
22. The acid H_2S reacts with an active metal, M. Which combination shows the correct role of H_2S , and product formed from the reaction?

	Role of H_2S	Product from H_2S reaction
A.	Oxidizing agent	$H_2(g)$
B.	Oxidizing agent	$S(s)$
C.	Reducing agent	$H_2(g)$
D.	Reducing agent	$S(s)$

23. What occurs during the operation of the voltaic cell $\text{Cu} \mid \text{Cu}^{2+} \parallel \text{Ag}^+ \mid \text{Ag}$?

- I. The blue colour of the $\text{Cu}^{2+}(\text{aq})$ solution will fade.
- II. $\text{NO}_3^-(\text{aq})$ ions in the salt bridge will migrate toward the copper electrode.
- III. The mass of the silver electrode will increase.

- A. I and II only
- B. I and III only
- C. II and III only
- D. I, II and III

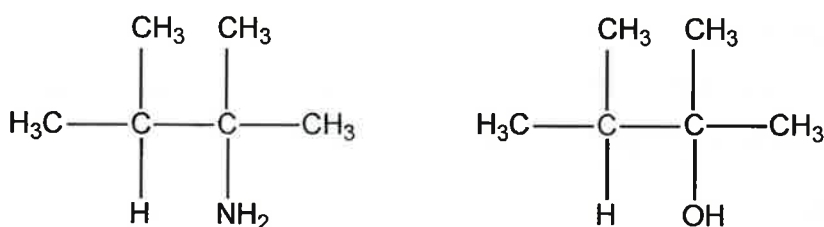
24. Which species will most readily react with propane?

- A. Br atom
- B. Br_2 molecule
- C. Br^- ion
- D. Br^+ ion

25. Which pair of compounds are structural isomers?

	1st compound	2nd compound
A.	$\text{NH}_2\text{CH}_2\text{COOH}$	$\text{NH}_2\text{CH}_2\text{OCH}_3$
B.	$\text{CH}_3\text{CH}_2\text{OH}$	CH_3COCH_3
C.	$\text{CH}_3\text{CH}_2\text{OH}$	CH_3OCH_3
D.	$\text{NH}_2\text{CH}_2\text{COOH}$	$\text{NH}_2\text{CH}_2\text{CH}_2\text{COOH}$

26. What is the correct classification for the two compounds given?



	Type of amine	Type of alcohol
A.	Primary	Primary
B.	Tertiary	Tertiary
C.	Tertiary	Primary
D.	Primary	Tertiary

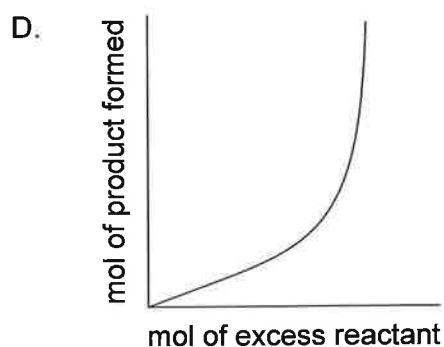
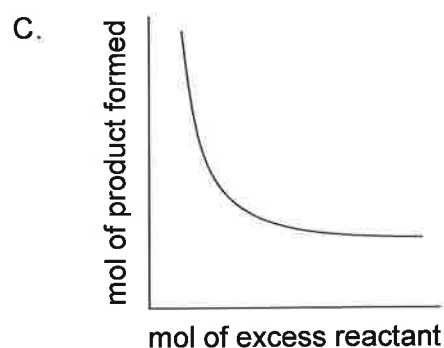
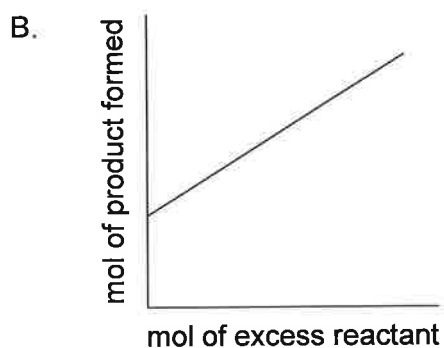
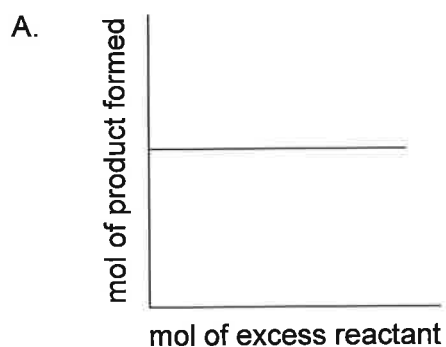
27. Which is the first product of distillation from the reaction of propan-1-ol with acidified potassium dichromate (VI)?

- A. CH_3COCH_3
- B. $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$
- C. $\text{CH}_3\text{CH}_2\text{CHO}$
- D. $\text{CH}_3\text{CH}_2\text{COOH}$

28. Which procedure is most likely to produce a systematic error in determining the original concentration of NaOH(aq) by titration with HCl(aq)?

- A. Repeating the titration only once instead of five times
- B. Using various burettes for each trial instead of the same one
- C. Using a varying number of drops of the indicator for the titrations
- D. Titrating the sample two days after preparing it instead of on the day it was prepared

29. Which graph shows the relationship between quantity of product formed and quantity of excess reactant after the limiting reactant is consumed?



30. Which type of data can be used to determine an empirical formula?

- A. Percentage composition
 - B. Enthalpy of combustion
 - C. ^1H NMR
 - D. Infrared spectroscopy (IR)
-